

1. Complete the following table.
(All particles are neutral unless otherwise indicated)

Isotope Name	Atomic Number	Mass Number	Nuclear Symbol	Number of Protons	Number of Neutrons	Number of Electrons
1. Carbon-14	6	14	$^{14}_6\text{C}$	6	8	6
2. Oxygen-16	8	16	$^{16}_8\text{O}$	8	8	8
3. Polonium-212	84	212	$^{212}_{84}\text{Po}$	84	128	84
4. Uranium-238	92	238	$^{238}_{92}\text{U}$	92	146	92
5. Hydrogen-2	1	2	^2_1H	1	1	1
6. Gold-197	79	197	$^{197}_{79}\text{Au}^{3+}$	79	118	76
7. Thorium-232	90	232	$^{232}_{90}\text{Th}$	90	142	90
8. Iodine-127	53	127	$^{127}_{53}\text{I}^-$	53	74	54
9. Lawrencium-257	103	257	$^{257}_{103}\text{Lr}$	103	154	103
10. Hydrogen-1	1	1	^1_1H	1	0	1
11. Nitrogen-14	7	14	$^{14}_7\text{N}^{3-}$	7	7	10
12. Helium-4	2	4	^4_2He	2	2	2
13. Calcium-40	20	40	$^{40}_{20}\text{Ca}^{2+}$	20	20	18
14. Carbon-12	6	12	$^{12}_6\text{C}$	6	6	6